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Lecture 16 . Thermodynamics II .

Worksheet . 1) A 85.2 g copper bar was heated to 221.32 . o. C and placed in a coffee cup calorimeter containing 425.0 mL of water at 22.55 . o. C. The final temperature of the water was recorded to be 26.15 . o. C. a. How much heat was

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gained by the water? b. How much heat was lost by the copper? c. What is the specific heat of copper?

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Coffee Cup Calorimetry • The reactants and products are the system. • The water which they are dissolved in and

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calorimeter make up the surroundings.

### **Thermodynamics II - AP Chemistry**

(VW, S & B: 2.3) D. Equilibrium: States of a system are most conveniently described when the system is in equilibrium, i. e. it is in steady-state. Often we will consider processes that change “slowly” -- termed quasi-steady.



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(VW, S & B: 2.3-2.4) Force Pressure Area

Gas 1 T1 Gas 2 T2 Gas 1 T3 Gas 2 T3

Wait copper boundary thermally

insulated

### **THERMODYNAMICS: COURSE INTRODUCTION**

Section 16 2 Heat And Thermodynamics

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### **Section 16 2 Heat And Thermodynamics Worksheet Answer**

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answers: Unit 10 Acids & Bases.

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## **Thermodynamics Worksheet Answer Key**

Objective: Write the 4 Laws of Thermodynamics. Define, explain and draw an example of entropy. Activities:

1. Students create a 3D pyramid on the laws of Thermodynamics.
2. Students

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view an internet video to learn more about entropy. 3. Students answer questions and create diagrams showing entr

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Lecture 14 Thermodynamics II  
Worksheet Key 1) Do the following  
processes produce an increase

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(23-27) The top of a 3/16 inch thick  
acrylic testing box is covered with an

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unknown  $\frac{1}{2}$  inch insulation material (black). The dimensions of box are 10 inch x 10 inch on each side. ... Principles of Engineering Activity 1.3.3  
Thermodynamics Answer Key - Page 5.  
Title: Activity 1.3.3 Thermodynamics Answer Key Subject: POE - Unit 1 - Lesson 1.3 ...

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#### **Activity 1.3.3 Thermodynamics**

#### **Answer Key**

Ch 1 - Introduction: Basic Concepts of Thermodynamics; Ch 2 - Properties of Pure Substances; Ch 3 - Heat Effects; Ch 4 - The First Law of Thermodynamics: Closed Systems; Ch 5 - The First Law of Thermodynamics: Open Systems; Ch 6 - The Second Law of Thermodynamics; Ch



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7 - Entropy; Ch 8 - Thermodynamics of Flow Processes; Ch 9 - Power Systems

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everyday life whether we realize it or not. In this lecture, we use thermodynamics to explain some basic observations made when cooking. Chemistry is also essential within living organisms, and we hear from researcher Lourdes Aleman about the importance of weak interactions known as hydrogen bonds.

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forward Energy As well as this, the  
Section of Heat and Thermodynamics  
worksheet answer key will include

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sections such as the Kinematics (linear acceleration) and Instantaneous Equilibrium (slowing down).

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